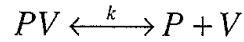


I. 閱讀以下文章，請於括號內之候選答案裡挑選出一個最正確的答案(每題 3 分)，遇底線處請填入正確答案(每題 5 分)。最後簡答文末所列問題，必要時列出算式，中英文皆可 (70%)。

※ 注意：請於試卷內之「非選擇題作答區」作答，並應註明作答之題號。

抗疫先鋒台大生機系陳教授開發一種可以一對一吸附病毒(V)的奈米顆粒(P)，它跟病毒之間的結合有如下的關係：



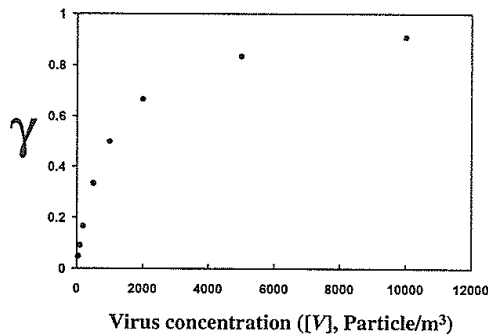
$$k = \frac{[P][V]}{[PV]}$$

此處， $k$  應為 [ (1) association; bonding; dissociation ] constant， $k$  的單位為 [ (2) concentration; mass; temperature; pressure; volume ]， $k$  與兩者間的 affinity 呈 [ (3) insignificant; negative; positive; significant ] correlation。

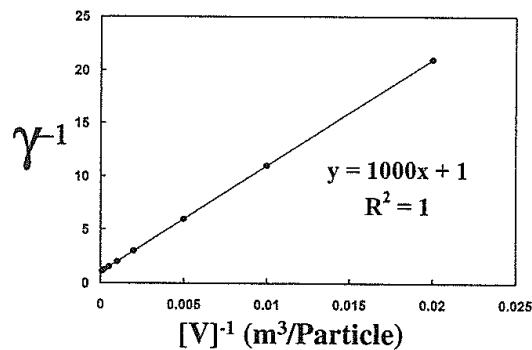
英明的陳教授於是將此神奇奈米顆粒噴在口罩上形成抗菌薄膜，並預測此抗菌薄膜與薄膜附近的病毒濃度合乎 Langmuir adsorption isotherm 的關係：

$$\gamma = \frac{[PV]}{[P]+[PV]} = \frac{1}{\frac{[P]}{[PV]}+1} = \frac{1}{\frac{k}{[V]}+1}$$

此處， $\gamma$  是口罩上的奈米顆粒與病毒結合的比例。A 同學分析不同口罩上的  $\gamma$  與口罩旁的病毒濃度 (Particle/m<sup>3</sup>)，得到如下圖關係。



由圖可知，口罩若長期暴露在一定濃度的病毒下， $\gamma$  值會 [ (4) decrease; hold constant; increase ] 到 [ (5) deprivation; threshold; saturation ]。為了證明陳教授 Langmuir adsorption isotherm 的推測，B 同學則用  $\gamma$  倒數與病毒濃度倒數作圖，得到完美的直線。



見背面

依上圖可得知，此奈米顆粒的  $k$  值為 (6) ，單位為 (7) 。

醫院的抗疫英雄戴著此口罩，發現  $\gamma$  值會隨時間( $t$ )呈指數上升。

$$\gamma = 1 - e^{-\frac{t}{\tau}}$$

在醫院使用 24 小時之後， $\gamma$  值變成 0.2，此時口罩附近的病毒濃度為 (8) (Particle/m<sup>3</sup>)。

請定性或定量回答以下問題 (每題 8 分)

- (9) B 同學為何比 A 同學更能證明陳大教授的推測？是根據什麼參數？
- (10) 試闡述  $\gamma$  值可能的量測方式？
- (11) 以  $\gamma$  值推估此抗疫口罩的 time constant 為多少？
- (12) 根據人體免疫數據顯示，人類免疫系統只能處理 100 Particle/m<sup>3</sup> 的病毒量，如此一來，請計算該口罩於醫院的有效使用時間？
- (13) 承上題，若口罩的 time constant 一定，而要開發能配戴 24 小時的防疫口罩，所使用奈米顆粒的  $k$  值應該為何？

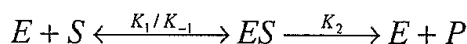
接次頁

II. 專業英文填空：請參考下列答案庫，限用英文，每題 2 分、文法錯誤扣 1 分 (30%)。

※ 注意：請於試卷內之「非選擇題作答區」作答，並應註明作答之題號。

答案庫(含單數名詞、原型動詞、形容詞與副詞)：activation; active; catalyst; catalyze; competitive; enzyme; first order; higher; intersect; kinetic; logarithmic; lower; gas constant; Michaelis constant; non-competitive; reciprocal; selective; second order; slope; specificity; steady; substrate; thermal; transient; uncompetitive.

Biological reactions \_\_\_\_\_ (14) \_\_\_\_\_ by enzymes are generally very \_\_\_\_\_ (15) \_\_\_\_\_ to certain chemicals known as the \_\_\_\_\_ (16) \_\_\_\_\_ of the enzymes; the chemicals bind to the \_\_\_\_\_ (17) \_\_\_\_\_ site of the enzyme to become an ES complex with a much lower \_\_\_\_\_ (18) \_\_\_\_\_ energy.



The reaction therefore can occur under mild conditions (e.g. temperature and pressure) with sufficient efficiency. Based on Michaelis-Menten model, the reaction rate is related to the concentration of ES complex as the following \_\_\_\_\_ (19) \_\_\_\_\_ kinetics.

$$V = -\frac{d[S]}{dt} = \frac{d[P]}{dt} = K_2[ES]$$

In the \_\_\_\_\_ (20) \_\_\_\_\_ state, the concentration of ES complex can be solved as follows.

$$\begin{aligned} \frac{d[ES]}{dt} &= 0 \\ K_1[E][S] &= (K_{-1} + K_2)[ES] \\ \frac{[ES]}{[E]} &= \frac{K_1[S]}{K_{-1} + K_2} = \frac{[S]}{K_M} \\ \frac{[ES]}{[E] + [ES]} &= \frac{[ES]}{[E]_t} = \frac{[S]}{K_M + [S]} \\ [ES] &= \frac{[S]}{K_M + [S]}[E]_t \end{aligned}$$

An enzymatic reaction rate therefore increases non-linearly with the substrate concentration and reaches a maximum as the substrate concentration being much higher than \_\_\_\_\_ (21) \_\_\_\_\_ .

$$V = K_2[ES] = K_2[E]_t \frac{[S]}{K_M + [S]} = V_{\max} \frac{[S]}{K_M + [S]}$$

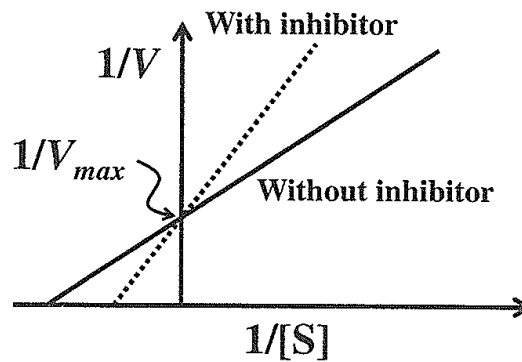
Since the kinetics is not linear, it will be more convenient to analyze the parameters by plotting the \_\_\_\_\_ (22) \_\_\_\_\_ values of the reaction rate against those of the substrate concentration.

見背面

The (23) of the plot will be proportional to the Michaelis constant.

$$\frac{1}{V} = \frac{1}{V_{max}} \left( \frac{K_M + [S]}{[S]} \right) = \frac{1}{V_{max}} + \left( \frac{K_M}{V_{max}} \right) \frac{1}{[S]}$$

The following plots show the existence of (24) inhibitors which will alter the apparent  $K_M$  (the slope of the plot) of the enzyme kinetics without changing the  $V_{max}$ . The higher the inhibitor concentration, the (25) the apparent  $K_M$ .



According to Arrhenius expression, the rate constant ( $k_2$ ) of an enzyme will be affected by the temperature ( $T$ ) and the activation energy ( $Ea$ ). The reaction for the substrate possesses an  $Ea$  (26) than other chemicals, which is the origin of the substrate (27) of the enzyme.

$$k_2 = Ae^{-Ea/RT}$$

$$V = k_2[ES] = Ae^{-Ea/RT} [ES]$$

$$\ln V = \ln A[ES] - \frac{Ea}{R} \left( \frac{1}{T} \right)$$

The best way to calculate the activation energy is to multiply the slope of the semi-logarithmic plot of  $V$  against  $(1/T)$  with the negative value of (28).